How do you work out the number of electrons in the outside shell?

Proton:

Neutron:

Electron:

Isotope:

**Group 7**

Forms \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ charged ions. By \_\_\_\_\_\_\_\_\_\_\_\_\_ an electron. This is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. Group 7 gets \_\_\_\_\_\_\_ reactive as you go down the group.

Description of Halogens (F,Cl, Br):

Symbol equation of Group 7 & Group 1:

Describe the displacement reactions:

**Group 1**

Forms \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ charged ions. By \_\_\_\_\_\_\_\_\_\_\_\_\_ an electron. This is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. Group 1 gets \_\_\_\_\_\_\_ reactive as you go down the group because \_\_\_\_\_\_\_\_\_ electron shells. The melting points \_\_\_\_\_\_\_\_\_\_ as you go down the group due to \_\_\_\_\_\_\_\_\_\_\_ metallic bonds as the number of electron shells \_\_\_\_\_\_\_\_\_\_\_\_\_.

Giant ionic lattice:

Symbol equation of Group 1 & water

Method for Flame Tests



**Transition Metal**

Describe colours of transition metals:

Describe uses of transition metals:

Symbol equation of thermal decomposition of transition metals:

Describe how to test for transition metal ions:

|  |  |  |
| --- | --- | --- |
| **Ionic Bonding** | **Covalent Bonding** | **Metallic Bonding** |
| DefinitionForms between which sorts of atoms?Draw a diagram of the bonding: NaCl | DefinitionForms between which sorts of atoms?Draw a diagram of the bonding: Cl2 Molecule | DefinitionForms between which sorts of atoms?Draw a diagram of the bonding |

Purifying Water: Testing with Barium Salts

Scientists & Periodic Table

* Dobereiner
* Newlands
* Mendeleev

Scientists & Structure of the Atom:

* Dalton
* JJ Thomson
* Rutherford
* Bohr